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PHYSICAL CONSTANTS OF INORGANIC COMPOUNDS (Continued)

No.	Name	Synonyms and Formulas	Mol. wt.	Crystalline form, properties and index of refraction	Density or spec. gravity	Melting point, °C	Boiling point, °C	Solubility, in grams per 100 cc: Cold water	Hot water	Other solvents
	Barium									
b45	fluosilicate	$BaSiF_6$	279.42	rhomb need	4.29^{16}_{4}	d 300		0.0768	0.09	i al; sl s a, NH
b46	formate	$Ba(CHO_2)_2$	227.38	col, rhomb, 1.573, 1.597, 1.636	3.21	d		27.7	39.7	i al, eth
b47	gluconate	$Ba(C_6H_{11}O_7)_2.3H_2O$	561.69	pr or rhomb leaf		-3 H_2O, (dr); d 170		3.3		i al
b48	hydride	BaH_2	139.36	gray cr	4.21	d 675	1400(?)	d to $Ba(OH)_2$ + H_2		d a
b49	hydroxide	$Ba(OH)_2.8H_2O$	315.46	col, monocl, 1.471, 1.501, 1.50	2.18	78	-H_2O, 78	5.6	94.7	sl s al; i acet
b50	hyponitrite	$BaN_2O_2.4H_2O$	260.41	wh cr powd	2.742					
b51	iodate	$Ba(IO_3)_2$	487.15	monocl	4.998	d		0.046	197	s HNO3, HCl
b52	iodate, hydrate	$Ba(IO_3)_2.H_2O$	505.17	col, monocl	4.8576	-H_2O, 200		v sl s	sl s	s HNO3, HCl; i al, acet, H
b53	iodide	BaI_2	391.15	col cr	5.15^{4}_{4}	740		170		sl 77
b54	iodide, hydrate	$BaI_2.2H_2O$	427.18	col rhomb, deliq	5.15	-H_2O, 88.9; -$2H_2O$, 539; d 740	2100	260		1.07 al; s acet
b55	iodide, hydrate	$BaI_2.6H_2O$	499.24	col, hex		15.7		410	v s	v s al
b56	laurate	$Ba(C_{12}H_{23}O_2)_2$	535.97	wh leaf cr		190		0.038	0.011	0.008 al; 0.009 eth
b57	lactate	$BaC_6H_{10}O_6$	299.41					0.883	1.04	
b58	malonate	$BaC_3H_2O_4.H_2O$	257.40	col				0.143	0.329	
b59	manganate	$BaMnO_4$	256.26	gray-grn, hex	4.85			v sl s		s a
b60	permanganate	$Ba(MnO_4)_2$	375.21	brnish cr	3.77	200		62.5	75.4	d al
b61	methylsulfate	$Ba(CH_3SO_4)_2.2H_2O$	395.56	col diff cr				s		s al
b62	molybdate	$BaMoO_4$	297.28	wh powd	4.65	1460		0.0356		sl s a
b63	myristate	$Ba(C_{14}H_{27}O_2)_2$	592.06					0.007	0.010	0.0006 al; 0.0014 eth; 0.046 s MeOH
b64	nitrate	Nitrobarite, $Ba(NO_3)_2$	261.35	col cub, 1.572	3.24	592	d	8.7	34.2	i al; sl s a
b65	nitride	Ba_3N_2	440.02	yel-br	4.783^{16}_{4}	1300 vac	d	d		
b66	nitrite	$Ba(NO_2)_2$	229.35	col, hex	3.23	d 217		67.5	300	sl s al
b67	nitrite, hydrate	$Ba(NO_2)_2.H_2O$	247.37	colorless, hex	3.172	d 115		83	109.6	1.6 al; v s HCl; i acet
b68	oxalate	BaC_2O_4	225.36	cr	2.458	d 400		0.0063	0.0228	i al; s NH4Cl
b69	oxide	BaO	153.34	col, cub, wh-yellsh powd, 1.98	5.72	1918	2000	3.48	90.8	s dil a, al; i acet, NH3
b70	oxide, perox	BaO_2	169.34	wh-gray powd	4.96	450	-O_2, 800	v sl s	d	s dil a; i acet
b71	oxide, perox, hydrate	$BaO_2.8H_2O$	313.46	col, hex	2.202	-$8H_2O$, 100		0.168	d	s dil a; i al, acet
b72	palmitate	$Ba(C_{16}H_{31}O_2)_2$	616.19	wh cr powd		d		0.0318	0.0078	0.0008 al; 0.0018 eth
b73	hypophosphate	BaP_2O_6	216.31	monocl				sl s		s a; v sl s al
b74	orthophosphate di-	$BaHPO_4$	233.33	wh, rhomb, 1.635, 1.637	4.165	d 410		0.01-0.02		s a, NH_4Cl
b75	orthophosphate, mono-	$Ba(H_2PO_4)_2$	331.31	triel	3.94			d	d	s a
b76	orthophosphate, tri-	$Ba_3(PO_4)_2$	601.96	wh, cub	4.46			i	i	s a
b77	pyrophosphate	$Ba_2P_2O_7$	448.62	wh, rhomb	3.98			0.01	sl s	s a, NH3 sl al
b78	hypophosphite	$Ba(H_2PO_2)_2.H_2O$	265.33	wh, monocl	2.90^{11}_{4}	d 100-150		30	33	i al
b79	propionate	$Ba(C_3H_5O_2)_2.H_2O$	301.50	rhomb, 1.518		d 300		48	67.9	0.06 al
b80	salicylate	$Ba(C_7H_5O_3)_2.H_2O$	437.63	wh need				s		
b81	selenate	$BaSeO_4$	280.30	wh, rhomb	4.75	d		0.016	0.138	s HCl; i HNO3
b82	selenide	$BaSe$	216.30	wh cub diss, mp, 2.268	5.02			d	d	d HCl
b83	metasilicate	$BaSiO_3$	213.42	col, rhomb, 1.573, 1.574, 1.576	4.399	1604		i	d	s HCl
b84	metasilicate, hydrate	$BaSiO_3.6H_2O$	331.52	rhomb, 1.543, 1.548, 1.548	2.59			0.17	d	
b85	stearate	$Ba(C_{18}H_{35}O_2)_2$	704.13	wh powd				0.004	0.006	0.003 al; 0.003 al; 0.011 eth
b86	succinate	$BaC_4H_4O_4$	253.37	wh powd				0.471	0.237	sl s al
b87	sulfate	Nat. barite, prec. blanc fixe, $BaSO_4$	233.40	wh, rhomb (monocl), 1.637, 1.638, 1.649	4.50	1580	tr 1149 monocl	0.0008	0.0033	0.00 s 3 N; 0.0008 s al s H_2SO_4
b88	persulfosulfate	$BaS_2O_6.4H_2O$	401.52	wh, monocl			d	52.2	d	d al
b89	sulfide, hydro-	$Ba(SH)_2.4H_2O$	275.50	yel, rhomb		d 60		s		i al
b90	sulfide, mono-	BaS	169.40	col, cub, mp 2.155	4.25	1900		d	d	i al
b91	sulfide, tetra-	$BaS_4.H_2O$	283.61	red or yel, rhomb	2.984	d 300		4.1	v s	i al, CS2
b92	sulfide, tri-	BaS_3	233.53	yel-grn cr		d 554		s	s	

No.	Name	Synonyms and Formulas	Mol. wt.	Crystalline form, properties and index of refraction	Density or spec. gravity	Melting point, °C	Boiling point, °C	Solubility, in grams per 100 cc: Cold water	Hot water	Other solvents
	Barium									
b02	sulfite	$BaSO_3$	217.40	col, sub hex		d		0.02[18]	0.002[18]	v s HCl
b03	tartrate	$BaC_4H_4O_6 \cdot H_2O$	303.52	wh cr	2.98[18]			0.026[18]	0.088[18]	0.032[18] al
b04	tellurite	$BaTeO_3 \cdot 3H_2O$	383.07	volum wh	4.2[18]	d >200		sl s	sl s	s HCl, HNO_3
b05	pyrotellurite, hydrogen	$Ba(HTeO_3)_2 \cdot H_2O$	391.76	volum ppt; hott yel; solid: wh				s	s	s s
b07	telluride	BaTe	264.94	yel-wh, sub, diss, mp 2,440	5.13					d s
b08	thiocarbonate	$BaCS_3$	245.54	yel, hex		d		0.08[18]	d	s al
b09	thiocyanate	$Ba(SCN)_2 \cdot 3H_2O$	289.53	needl, deliq	2.28[18]	$-H_2O$, 150		0.3[18]	s	35[18] al
b100	thiosulfate	BaS_2O_3	249.47	wh, rhomb		d 220		0.2		
b101	thiosulfate, hydrate	$BaS_2O_3 \cdot H_2O$	267.48	wh, rhomb	3.5[18]	d 100		0.203[18]		i al
b102	titanate	$BaTiO_3$	233.24	tetr and hex, 2.40		tetr 1,017, hex 1,806				
b103	tungstate	$BaWO_4$	385.19	col, tetr	5.04			sl s	sl s	d s
b104	pyrovanadate	$Ba_2V_2O_7$	488.66	wh cr		963				
b105	Beryllium	Glucinum, Be	9.01218	grey met, hex	1.85[18]	1278 ± 5	2970[760 mm]	s	sl s, d	s dil a, alk; i Hg
b106	acetate	$Be(C_2H_3O_2)_2$	127.10	col pl		d 300				i al, eth, CCl_4
b107	acetate, basic	$BeO(C_2H_3O_2)_4$	406.32	oct	1.36	284	331	sl d	d	s chl, as al; sl s al, eth
b108	acetate propionate, basic	$Be_4O(C_2H_3O_2)_2(C_3H_5O_2)_4$	448.40			127	330			
b109	aluminate	Nat. chrysoberyl, $BeAl_2O_4$	126.97	rhomb, 1.747, 1.748, 1.757	3.78	1870				i a
b110	aluminum silicate	Nat. beryl, $Be_3Al_2(SiO_3)_6$	537.51	transp, hex, col, 1.580, 1.587	2.66	1410 ± 100				i a
b111	aluminum silicate	Nat. euclase, $BeAl(SiO_4)(OH)$	200.17	monocl, 1.652, 1.655, 1.671	3.1					
b112	benzenesulfonate	$Be(C_6H_5O_3S)_2$	323.35	monocl				v s	v s	v s al, acet, as al; i eth, bz, CH_4, CCl_4
b113	orthoborate, basic	Nat. hambergite, $Be_2(OH)BO_3$	93.84	rhomb, 1.560, 1.591, 1.631	2.37					
b114	bromide	$BeBr_2$	168.83	wh needl, deliq	3.465[18]	490 ± 10 subl	520	s	v s	s al, eth; i bz; 18.56 pyr
b115	butyrate, basic	$Be_4O(C_4H_7O_2)_6$	574.64				239[10]			
b116	carbide	Be_2C	30.04	yel, hex	1.90[18]	>2100 d		d	d	s a; d alk
b117	carbonate, basic	$BeCO_3 + Be(OH)_2$	112.05	wh powd				i	d	s a, alk
b118	chloride	BeCl	79.92	col needl, deliq	1.899[18]	405	520 (438)	v s	v s, d	v s al, eth, pyr; sl s bz, chl, CS_2; i NH_3, acet
b119	fluoride	BeF	47.01	col, amorph, <1.33	1.986[18]	subl 800		s	s	sl s alt; s H_2SO_4
b120	hydride	BeH_2	11.03	wh cr		d 120		d	d	i eth, tol
b121	iodide	BeI_2	262.82	col needl	4.325[18]	510 ± 10	590	d	d	s al, eth, CS_2
b122	nitrate	$Be(NO_3)_2 \cdot 3H_2O$	187.07	wh-yel cr, deliq	1.56-7	60	142	v s	v s	v s al
b123	nitride	Be_3N_2	55.05	col, cub		2200 ± 100	d 2240	d	d	d a, some alk; i al
b124	oxalate	$BeC_2O_4 \cdot 3H_2O$	151.08	rhomb, n 1.477		$-H_2O$, 100; $-3H_2O$, 220	d 380	38.22[18]		
b125	oxide	Nat. bromellite, BeO	25.01	wh hex, 1.719, 1.733	3.01	2530 ± 30	bp 3900	0.00072[18]		s conc H_2SO_4, fus KOH
b126	oxide	$BeO \cdot xH_2O$		wh amorph powd or gel		d		i	i	s a, alk, $(NH_4)_2CO_3$
b127	2,4-pentanedione deriv	Beryllium acetylacetonate $Be(C_5H_7O_2)_2$	207.23	wh, monocl	1.1[18]	108	270	sl s	d	s al, eth, a
b128	orthophosphate	$Be_3(PO_4)_2 \cdot 3H_2O$	271.03			$-H_2O$, 100		s	s	s as a
b129	propionate basic	$Be_4O(C_3H_5O_2)_6$	490.43			120				
b130	selenate	$BeSeO_4 \cdot 4H_2O$	224.03	col, rhomb, 1.466, 1.501, 1.503	2.03	$-2H_2O$, 100; $-4H_2O$, 300		56.7[18]	s	
b131	di-silicate	Nat. bertrandite, $Be_4Si_2O_7(OH)_2$	338.33	rhomb, 1.591, 1.605, 1.604	2.4					
b132	orthosilicate	Nat. phenakite, Be_2SiO_4	110.11	trigl, col, 1.654, 1.670	3.0					i a
b133	stearate (sol'l)	$Be(C_{18}H_{35}O_2)_2$	575.97	wh, waxy		48		i	i	i al; s eth, CCl_4
b134	sulfate	$BeSO_4$	105.07		2.443	d 550-600		i	d to $BeSO_4 \cdot 4H_2O$	
b135	sulfate, hydrate	$BeSO_4 \cdot 4H_2O$	177.14	col, tetr, 1.472, 1.440	1.713[18]	$-2H_2O$, 100	$-4H_2O$, d>[illegible]	42.5[18]	100[18]	sl s conc H_2SO_4; i al
b136	sulfide	BeS	41.08	reg	2.16			d	d	
b137	Bismuth	Bi	208.9804	rhomb silver-wh or reddish met	9.80	271.3	1560 ± 5[760]	i	i	s h H_2SO_4, HNO_3, aq reg; sl s h HCl
b138	acetate	$Bi(C_2H_3O_2)_3$	386.12	wh cr		d		i	i	s as a

Name	Synonym and Formula	Mol. wt.	Crystalline form, properties and index of refraction	Density or sp. gr. gravity	Melting point, °C	Boiling point, °C	Solubility, in grams per 100 cc: Cold water	Hot water	Other solvents
Cadmium									
pyrophosphate	$Cd_2P_2O_7$	398.71	wh or leaf	4.06[20]	above red heat		sl s	s	s a, NH_3
phosphate, dihydrate	$Cd(H_2PO_4)_2.2H_2O$	312.41	col, tricl	2.71[15]	sl 100				l al, eth; s HCl
phosphide	Cd_3P_2	399.13	grn, tetr need	5.00	700				sol HCl; s esp conc HNO_3
potassium cyanide	$Cd(CN)_2.2KCN$	294.68	col, glassy, oct	1.847			33.3	100	l al
potassium sulfate	$CdK_2(SO_4)_2.2H_2O$	322.69	tricl col tab	2.022[16]			13.8[16]	17.4[100]	
salicylate	$Cd(C_7H_5O_3)_2.H_2O$	404.65	wh need				sl s	s	s al, eth, glycerol; s, NH_4OH
selenate	$CdSeO_4.2H_2O$	291.39	rhomb	3.83	−H_2O, 100; −$2H_2O$, 170		v s		
selenide	CdSe	191.36	grnish or red powd, hex	5.81[0]	>1350		i	i	s a
selenite	$CdSeO_3$	188.38	col, rhomb, 1.739	4.01	1212		v sl s		
sulfate	$CdSO_4$	208.40	wh, rhomb	4.691[20]	1000		75.5	63.8[100]	l al, acet, NH_3
sulfate, hydrate	$CdSO_4.H_2O$	226.45	monocl	3.79[20]	tr 108		s	s	l al
sulfate, hydrate	$CdSO_4.7H_2O$	334.57	col, monocl	2.48	tr 4		s		l al
sulfate, hydrate	$3CdSO_4.8H_2O$	769.50	col, monocl, 1.564	3.09	tr 41.5		11.[illegible]	s	
sulfide	Nat. greenockite, CdS	144.46	yel-or, hex, 2.506, 2.529	4.82	1750[press]	subl in N_2 980	0.0[illegible]13[18]	colloid	s a; v sl s NH_4OH
sulfite	$CdSO_3$	192.46	cr		d		sl s		l al; s a, NH_4OH
tartrate	$CdC_4H_4O_6$	260.47	wh cr powd				sl s		s a, NH_4OH
telluride	CdTe	240.00	blk, sub	6.20[4]	1041		i		l al; d HNO_3
tungstate	$CdWO_4$	360.25	yel cr				0.0[illegible]		s NH_4OH
Cadmium complexes									
tetramminecadmium perhenate	$[Cd(NH_3)_4](ReO_4)_2$	681.16		3.714[0]					0.037 conc NH_4OH
tetrapyridine rutheniumfluosilicate	$[Cd(C_5H_5N)_4]SiF_6$	570.89	wh, tricl	2.282					
Calcium	Ca	40.08	silv wh soft met, sub	1.54	839±2	1484	d to H_2 + $Ca(OH)_2$	d	s a, liq NH_3; sl s al; l bz
acetate	$Ca(C_2H_3O_2)_2$	158.17	col cr; 1.55, 1.56, 1.57		d		27.[illegible]	29.7[100]	sl s al
acetate, dihydrate	$Ca(C_2H_3O_2)_2.2H_2O$	194.21	col need		−H_2O, 54		34.[illegible][100]	33.8[100]	
acetate, monohydrate	$Ca(C_2H_3O_2)_2.H_2O$	176.19	col need		d		43.[illegible][0]	34.3[100]	sl s al
aluminate	$CaAl_2O_4$ (or $CaO.Al_2O_3$)	158.04	wh monocl, tricl or rhomb; 1.643, 1.648, 1.663	2.98[14]	1600		d		s HCl; l HNO_3, H_2SO_4
orthoaluminate	$Ca_3Al_2O_6$ (or $3CaO.Al_2O_3$)	270.20	wh, sub, 1.710	3.038[14]	d 1535		i		s a
orthoaluminate monohydrate	$3CaO.Al_2O_3.6H_2O$	378.29	col, oct, 1.603	2.52[20]	d 700–800		d		
dialuminosilicate	$2CaO.Al_2O_3.SiO_2$	274.30	col, tetr, 1.669, 1.658	3.018	1590±2				d a
dialuminosilicate	Nat. anorthite, $CaAl_2Si_2O_8$ (or $CaO.Al_2O_3.2SiO_2$)	278.21	wh, tricl, 1.5833	2.763	1551				
arsenate	$Ca_3(AsO_4)_2$	398.08	col amorph powd	3.620	1455		0.0[illegible][18]		
arsenate, trihydrate	Nat. haidingerite, $3CaO.As_2O_5.3H_2O$	396.04	col, rhomb, 1.590, 1.602, 1.638	2.967					
arsenide	Ca_3As_2	270.06	red cr	3.031[20]	d		d	d	d a; s h HNO_3
azide	$Ca(N_3)_2$	124.12	col, rhomb, hyg		−$3H_2O$, 110; exp 144–166		38.[illegible]	45[0]	0.211[16] al; l eth
borate	$Ca(C_2H_5O_2)_2.3H_2O$	336.36	col, rhomb	1.436	−$3H_2O$, 110		2.7[illegible]	8.3[100]	
metaborate	$Ca(BO_2)_2$	125.70	col, flat rhomb pr, 1.550, 1.660, 1.680		1154		sl s		s a, NH_4 salts; sl s as a
metaborate, hexahydrate	$Ca(BO_2)_2.6H_2O$	233.79	col, tetr, 1.520, 1.572	1.88			0.2[illegible][0]		
metaborate	CaB_2O_4	195.32	readily vitrified		986				
bromate	$CaBr_4$	161.95	blk, sub	2.3[0]	2235		i	i	s HNO_3; sl s conc H_2SO_4
bromide	CaBr	199.90	col, rhomb need, deliq	3.353[20]	sl d 730	806–812	142[0]	312[100]	s al, acet, a; sl s NH_3, MeOH
bromate	$Ca(BrO_3)_2.H_2O$	313.91	monocl cr	3.329	−H_2O, 110		v s	v s	
bromate, hexahydrate	$CaBr_2.6H_2O$	307.99	col, hex cr	2.295	38.2	149	594[0]	1360[100]	s al, acet, a

PHYSICAL CONSTANTS OF INORGANIC COMPOUNDS (Continued)

No.	Name	Synonyms and Formulas	Mol. wt.	Crystalline form, properties and index of refraction	Density at spec. gravity	Melting point, °C	Boiling point, °C	Solubility, in grams per 100 cc: Cold water	Hot water	Other solvents
	Chromium									
418	(III) orthophosphate	$CrPO_4.2H_2O$	183.00	vlt cr	2.4281			sl s		s s, alk; i ac s
417	(III) orthophosphate	$CrPO_4.6H_2O$	255.06	vlt, trcl, 1.568, 1.591, 1.599	2.1210	100				
418	pyrophosphate	$Cr_2(P_2O_7)_3$	729.81	pa grn, monocl	2.2			i	i	s alk
419	phosphide, mono-	CrP	82.97	gray-blk cr	5.70			i		s HNO_3, HF
420	silicide	Cr_3Si	212.17	gray, tetr pr	5.50			i	i	s HCl, HF; i H_2SO_4, HNO_3
421	(II) sulfate	$CrSO_4.7H_2O$	274.17	bl cr				12.054	d	s ls al; s NH_4OH
422	(III) sulfate	$Cr_2(SO_4)_3$	392.18	vlt or red powd	3.012			i, s		sl s al; i a
423	(III) sulfate	$Cr_2(SO_4)_3.15H_2O$	662.41	vlt, amorph sc	1.8670	100	~100 H_2O, 100	s	d 87	i al
424	(III) sulfate	$Cr_2(SO_4)_3.18H_2O$	716.45	bl vlt, rub oct, 1.544	1.70	~120 H_2O, 100		120 s	s	s al
425	(II) sulfide, mono-	CrS	84.06	blk powd, hex	4.85	1550		i		v s a
426	(III) sulfide, sesqui-	Cr_2S_3	200.18	brn-blk powd	3.77	~4, 1350		i, s	i	s HNO_3; d al
427	(III) sulfite	$Cr_2(SO_3)_3$	344.18	grnish-wh	2.2	d				
428	(II) tartrate	$CrC_4H_4O_6$	201.07	bl powd	2.33			i	i	sl s a; i ac s
429	Chromium compounds									
	hexammine chromium(III) chloride	$[Cr(NH_3)_6]Cl_3.H_2O$	275.55	yel cr	1.565			s		
430	hexaureachromium(III) fluosilicate	$[Cr(CON_2H_4)_6](SiF_6)_3.3H_2O$	1304.04	lt grn leaf				0.0320		s al
431	hexaureachromium(III) perrhenate	$[Cr(CON_2H_4)_6](ReO_4)_3$	1142.92	grn nedl	2.662			1.186		0.607 al
432	chloropentammine chromium chloride	$[Cr(NH_3)_5Cl]Cl_2$	243.51	red, oct	1.898			0.450		i HCl
433	Cobalt	Co	58.9332	silv gray met, cub	8.9	1495	2870	i	i	s a
434	(II) acetate	$Co(C_2H_3O_2)_2$	176.07	grn, oct		d 100		5 rdt solubility		s s, glac ac s
435	(II) acetate	$Co(C_2H_3O_2)_2.4H_2O$	249.08	red-vlt, monocl, delic, 1.542	1.7050	~41 H_2O, 140		s	s	s s, al
436	aluminate	(approx) Thenard's blue, $CoAl_2O_4$	176.90	bl, cub				i	i	
437	(II) orthoarsenate	$Co_3(AsO_4)_2.8H_2O$	598.75	vlt-red, monocl, 1.626, 1.661, 1.669	3.1781	d		i	i	s dil a, NH_4OH
438	arsenic sulfide	Nat. cobaltite, CoAsS	165.92	gray-wtish	6.2–6.3	d				
439	arsenide	CoAs	132.79	yel powd	8.28	960			i	s HCl, H_2SO_4; s HNO_3, aq reg
440	(II) benzoate	$Co(C_7H_5O_2)_2.4H_2O$	373.23	gray red leaf		~41 H_2O, 115		i s		
441	bromide, mono-	CoBr	58.74	pr	4.1810			d	d	s HNO_3, aq reg
442	(II) bromate	$Co(BrO_3)_2.6H_2O$	422.84	red, oct				15.50		s NH_4OH
443	(II) bromide	$CoBr_2$	218.75	grn, hex, delic	4.909	678 (in N_2)		55.7	58.1	77.1 al; 58.6 MeOH; s eth, acet
444	(II) bromide hexahydrate	$CoBr_2.6H_2O$	326.84	red-vlt pr, delic	2.46	47–48, ~41 H_2O, 100	~61 H_2O, 150	s red solut	153.3	s blk solut, al, s, eth
445	bromoplatinate	$CoPtBr_6.12H_2O$	949.66	trcl	2.712					
446	carbonate	Nat. spherocobaltite, $CoCO_3$	118.94	red, trig, 1.833, 1.81	4.18	d		i	i	s a; i NH_3
447	(II) carbonate, basic	$2CoCO_3.Co(OH)_2.H_2O$	534.74	vltred pr				i	d	s a, $(NH_4)_2CO_3$
448	carbonyl tetra-	Dicobalt octacarbonyl, $[Co(CO)_4]_2$ or $Co_2(CO)_8$	341.95	or cr or dk br, mixture	1.7106	51	d 52	i	i	sl s al; s CS_2, eth
449	carbonyl, tri-	Tetracobalt dodecacarbonyl, $[Co(CO)_3]_4$ or $Co_4(CO)_{12}$	571.86	blk cr				sl s		s bz; d HF
450	(II) chlorate	$Co(ClO_3)_2.6H_2O$	333.93	red, oct, delic, 1.55	1.97	50	d 100	558.3	v s	s al
451	(II) perchlorate	$Co(ClO_4)_2$	257.82	red nedl 1.510, 1.690	3.197			100	115	s al, acet
452	(II) perchlorate	$Co(ClO_4)_2.4H_2O$	347.88	red, hex		143		100.13	115.10	v s al, acet; i $CHCl_3$
453	perchlorate	$Co(ClO_4)_3.6H_2O$	365.93	red pr		d 1034	d	250		s al, acet
454	(III) perchlorate	$Co(ClO_4)_3.6H_2O$	365.93	red, oct, delic, 1.55		d 102		255		s al, acet
455	(II) chloride	$CoCl_2$	129.84	bl, hex, hygr	3.356	724 (in HCl gas)	1049	45	105	54.4 al; 8.6 acet; 38.5 MeOH; sl s eth
456	(III) chloride	$CoCl_3$	165.79	red or or yel cr	2.94	sub		s		

* Several chromic salts exist in two forms, a soluble and insoluble modification.